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GAS V IS THE'

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Combined Gas Law The Combined Gas  
Law combines Charles' Law, Boyle's Law  
and Gay Lussac's Law. The Combined  
Gas Law states that a gas' (pressure  $\times$   
volume)/temperature = constant. The  
combined law for gases. Example: A gas  
at 110kPa at 30.0°C fills a flexible  
container with an initial volume of 2.00L.

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The combined gas law combines the three gas laws: Boyle's Law, Charles' Law, and Gay-Lussac's Law. It states that the ratio of the product of pressure and volume and the absolute temperature of a gas is equal to a constant. When Avogadro's law is added to the combined gas law, the ideal gas law results. Unlike

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Answers: COMBINED GAS LAW

Remember to convert all temperatures to Kelvin. P 1 V 1 T 1 P 2 V 2 T 2 1 1.5 atm 3.0 L 20. C 293K 2.5 atm 1.9 L 30. C 303K 2 720 torr 256 mL 25 C 298 K 8.0x10<sup>2</sup> torr 250 mL 50. C 323 K 3 600. mmHg 2.5 L 22 C 295 K 760 mmHg 1.8 L 270 K 4 1.2 atm 750 mL 0.0 C 273.0 K 2.0 atm 500. mL 25 C

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Ideal And Combined Gas Law Chemfiesta Answers MIXED GAS LAWS WORKSHEET Created by Tara L. Moore at Directions: Answer each question below. Then write

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the name of the gas law used to solve each question in the left margin next to each question. 1. A gas occupies 3.5L at 2.5 mm Hg pressure.

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